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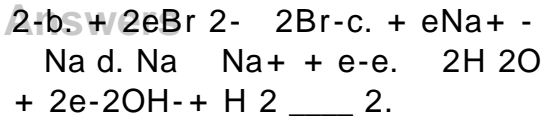
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Multiple Choice. Identify the
choice that best completes the
statement or answers the question.

____ 1. The half-reaction that
occurs at the cathode during the
electrolysis of molten sodium
bromide is _____. a. $+ 2e^- 2Br^- \rightarrow Br_2$

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Multiple Choice Answers 14.

Questions 14-17 The spontaneous reaction that occurs when the cell in the picture operates is as follows: $2\text{Ag}^+ + \text{Cd (s)} \rightarrow 2\text{Ag (s)} + \text{Cd}^{2+}$ (A) Voltage increases. (B) Voltage decreases but remains $>$ zero.

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$\text{Zn(s)} + \text{Ni}^{2+}(\text{aq}) \rightarrow \text{Ni(s)} + \text{Zn}^{2+}(\text{aq})$ (a) Identify M and M^{2+} in the diagram and specify the initial concentration for M^{2+} in solution. Electrons flow from the

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Answers to the cathode in a voltaic electrochemical cell. The anode is where oxidation occurs, and in the reaction above, Zn(s) is oxidized.

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 $\text{Sr}(s) + \text{Mg}^{2+} \rightleftharpoons \text{Sr} +$
 $\text{Mg}(s)$ Consider the reaction
represented above that occurs at

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25 °C. All reactants and products
are in their standard states.

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the cell potential and free energy
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Answer: (a) tin electrode is the
cathode; cathode is the site of
reduction (gain in electrons) and

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Answers
will convert metal ions into a
metal. (b) (see diagram) (c) red:
 $\text{Sn}^{2+} (\text{aq}) + 2 \text{e}^{-} \rightarrow \text{Sn} (\text{s}) \quad E^{\circ} =$
 -0.14 V oxid: $\text{X} (\text{s}) \rightarrow \text{X}^{3+} (\text{aq}) + 3 \text{e}^{-} \quad E^{\circ} =$
 $+0.74 \text{ V}$ $E^{\circ} \text{ cell} =$
 $+0.60 \text{ V}$ red: $\text{X}^{3+} (\text{aq}) + 3 \text{e}^{-} \rightarrow \text{X} (\text{s})$

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$\text{Sr}(s) + \text{Mg}^{2+} \rightleftharpoons \text{Sr} +$

$\text{Mg}(s)$ Consider the reaction

represented above that occurs at

25 ° C. All reactants and products

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Answers are in their standard states. The value of the equilibrium constant, K_{eq} , for the reaction is 4.2×10^{17} at 25°C .

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Answer: Electrochemistry
questions. This is the currently
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MAIN TYPES Of Electrochemical
Cells : 1. Galvanic (voltaic) cells –
which are thermodynamically
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(battery) 2. Electrolytic cells – which are thermodynamically unfavorable and require external e^- source (a direct current or DC power source)

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Electrochemistry. For the galvanic
cell described below, the correct

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line notation is: $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$ ($E^\circ = 1.36\text{V}$)
 $\text{Cu} + \text{e}^- \rightarrow \text{Cu}$ ($E^\circ = 0.52\text{V}$)
 $\text{Cu (s)} | \text{Cu}^+ (\text{aq}) || \text{Cl}_2 (\text{g}) | 2\text{Cl}^- (\text{aq}) | \text{Pt (s)}$
 $\text{Pt (s)} | \text{Cu (s)} | \text{Cu}^+ (\text{aq}) || \text{Cl}_2 (\text{g}) | 2\text{Cl}^- (\text{aq}) | \text{Pt (s)}$

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Answer the following questions
regarding the electrochemical cell
shown above. (a) Write the

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Answers balanced net-ionic equation for the spontaneous reaction that occurs as the cell operates, and determine the cell voltage. (b) In which direction do anions flow in the salt bridge as the cell operates? Justify your answer. (c) If 10.0 mL of 3.0-molar AgNO_3

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Creator of the award-winning chemistry website, www.adriandingleschemistrypages.com. The focus of Mr. Dingle ' s teaching career has been on preparing students for standardized tests; AP® and SAT® tests in the United States, GCSE ' s and A

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Answers in the United Kingdom, and International Baccalaureate in both countries. An Englishman, he lives in Atlanta, Georgia, where he teaches at The Westminster Schools. He holds a B.Sc. (Hons.) Chemistry, and a Postgraduate Certificate in Education, both from

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