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Answer the
following questions
in the space

provided. 1. Given
the following
equation: $C_3H_8(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the
value of the
coefficient x in this

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equation? 40.07
g/mol b. What is
the molar mass of
C₃H₄? 2 mol O
2:1 mol H₂O c.
What is the mole
ratio of O₂ to H

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required to solve
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problems when. ...

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molecules of
hydrogen peroxide
(in aqueous

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decompose to produce two molecules of liquid water and one molecule of oxygen gas. Chapter 9: Standard Review Worksheet Start studying Chapter 9: Stoichiometry Review. Learn vocabulary, terms, and more with

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SECTION 3
PROBLEMS Write
the answer on the
line to the left.
Show all your work

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in the space
provided. 1.
The actual yield of
a reaction is 22 g
and the theoretical
yield is 25 g.

Calculate the
percentage yield.

2. 6.0 mol of N₂
are mixed with
12.0 mol of H

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Theoretically, how many moles of NH_3 will be produced?

PROBLEMS Write

the answer on the line to the left,

Show all your work in the space

provided. 1 88%

The actual yield of a reaction is 22 g and the theoretical yield is 25 g.

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Calculate the
percentage yield.
2. 6.0 mol of N_2
are mixed with
12.0 mol of H_2
according to the ...

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~~the answer on the~~

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line to the left.

Show all your work
in the space

provided. 1. The
following equation
represents a

laboratory
preparation for
oxygen gas: ...

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...

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Standard Review Worksheet 1.

Answers will vary.

An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide

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(in aqueous solution)
decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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specifically acquire

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answers can be
one of the options
to accompany you
following having...

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Composition
Stoichiometry -
deals with mass
relationships of
elements in
compounds

Reaction
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Involves mass
relationships

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between reactants
and products in a
chemical reaction.

I. Reaction

Stoichiometry
Problems A. Four
problem Types,
One Common
Solution.

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SHORT ANSWER

Answer the
following questions
in the space

provided. 1. Given
the following

equation: $C_3H_8(g)$

$+ xO_2(g) \rightarrow 3CO$

$2(g) + 2H_2O(g)$ 4

a. What is the

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value of the
coefficient x in this
equation? 40.07
g/mol b. What is
the molar

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SHORT ANSWER
Answer the

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following questions
in the space
provided. 1. Given
the following
equation: $C_3H_4(g)$
 $+ xO_2(g) \rightarrow 3CO$
 $2(g) + 2H_2O(g)$ 4
a.

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