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2--Calculating Equilibrium Chapter 15

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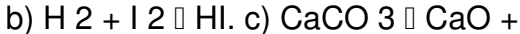


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## Chemical Equilibrium

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5. Amongst the following chemical reactions, the irreversible reaction is:



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## Chemical Equilibrium

The reaction is not reversible. 6. When the rate of forward reaction becomes equal to backward reaction, this state is termed as: a) Chemical equilibrium. b) Reversible state

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## Chemical Equilibrium

According to Le Chatelier's principle, increasing the pressure of the system will shift the equilibrium to the left (towards reactants).  $2\text{Na (s)} + 2\text{H}_2\text{O (l)} \rightleftharpoons \text{H}_2 \text{ (g)} + 2\text{NaOH (aq)}$

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## Chemical Equilibrium

Answers to Chemistry End of Chapter Exercises. 1. The reaction can proceed in both the forward and reverse directions. 3. When a system has reached equilibrium, no further changes in the reactant and product concentrations occur; the reactions continue to occur, but at equivalent

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~~13.1 Chemical Equilibria~~ □ Chemistry  
9.5: Chemical Equilibrium Conditions  
for Equilibrium and Types of  
Equilibrium. It may be tempting to think  
that once equilibrium has been  
reached,... Equilibrium Constant.

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Answer Consider the hypothetical reversible reaction in which reactants A and B react to form products C... Reaction Quotient. The ...

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FOR ALL EQUILIBRIUM PROBLEMS,

YOU MUST: 1) Write all equilibrium

equations 2) Write all equilibrium

concentrations 3) Write all equilibrium

expressions SET A: a) What is the

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## Chemical Equilibrium

~~Answers~~ Constant expression for the reaction:  $3 \text{Fe}(s) + 4 \text{H}_2\text{O}(g) \rightleftharpoons 3 \text{FeO}(s) + 4 \text{H}_2(g)$  b) The equilibrium constant,  $K_c$ , for the reaction:

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~~Worksheet Answer Keys~~

Calculate the equilibrium constant. 1.4



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## Chemical Equilibrium

$\text{ZnO(s)} + \text{CO(g)} \rightleftharpoons \text{Zn(s)} + \text{CO}_2\text{(g)}$ . At 400 K the  $K_c$ -value for this reaction is 2,78. If there is 5,6 g  $\text{CO(g)}$  in the 3 dm<sup>3</sup> container at equilibrium, calculate the concentration of the  $\text{CO}_2$ . 1.5 3,70 mol of A is placed in a 4 dm<sup>3</sup> container and heated. When equilibrium is

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## ~~1 Worksheet: Chemical equilibrium~~ UP

Here we have provided all the important Multiple choice questions and answers of Chapter Chemical Equilibrium of 10th class chemistry.  
MCQ 1: When diluting acid always add  
A. water to acid B. acid to water

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~~Chapter 1 Chemical Equilibrium  
Important MCQs with Answers ...~~

At equilibrium, the concentration of all reactants and products are determined by the initial concentrations and the equilibrium constant. No such conclusion can be made. g) False.

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~~Answers~~

1. If you have this equilibrium reaction:  
 $2\text{HI}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{I}_2(\text{g})$  What will  
happen to the equilibrium add more  $\text{I}_2$   
or oxygen ( $\text{O}_2$ )? 2. If you have a this

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## Chemical Equilibrium

Answers:  $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$ . There is .0220 M  $\text{PCl}_3$  that you start with. And then at equilibrium at 15. liters you have .0333  $\text{PCl}_5$ . So how many moles  $\text{pcl}_5$  are there? How many mole/liters are there for  $\text{pcl}_5$ ? How to do calculate  $K_c$  ...

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Equilibrium means that opposing processes are in balance. Reversible reactions balance each other because they take place at equal rates. This is called chemical equilibrium. Be mindful that equilibrium is a state of action;

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~~Answers~~ movement is constant.

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feel confident in answering the  
question in the exam & increases your

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Which of the following are equal for a



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Answer: A chemical system at equilibrium? If all are equal, answer E. a. the concentrations of reactant and products are equal b. the rate constants for the forward and reverse reactions are equal c. the time that a particular atom or molecule spends as a reactant and product are equal d.

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## ~~Big Picture Introductory Conceptual Questions~~

For the reaction...  $\text{heat} + \text{N}_2 + \text{O}_2 \rightleftharpoons 2\text{NO}$ . If the heat is removed to the chemical system, the equilibrium will shift \_\_\_\_\_. answer choices. left. right. both left and right. neither left nor

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right. Tags:

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The equilibrium constant for the reaction:  $2 \text{NO}(g) \rightleftharpoons \text{N}_2(g) + \text{O}_2(g)$  is  $2.60 \times 10^{-3}$  at  $1100^\circ \text{C}$ . If 0.820 mole of  $\text{NO}(g)$  and 0.223 mole each of  $\text{N}_2(g)$  and

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$\text{O}_2$  (g) are mixed in a 1.00 liter container at  $1100^\circ\text{C}$ , what are the concentrations of  $\text{NO}$  (g),  $\text{N}_2$ (g), and  $\text{O}_2$  (g) at equilibrium?

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The concept of chemical equilibrium

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## Chemical Equilibrium

was developed after Berthollet(1803) found that some chemical reactions are reversible. For any reaction mixture to exist at equilibrium, the rates of the forward and backward (reverse) reactions are equal.  $\alpha A + \beta B \rightleftharpoons \gamma S + \delta T$

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