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3.1 Modern Atomic Theory Atomic
Theory through the Nineteenth

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Century. The earliest recorded discussion of the basic structure of matter comes... Radioactivity. The second line of investigation began in 1896, when the French physicist Henri Becquerel (1852–1908)... The Atomic Model. Once scientists ...

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SECTION 3 Name Class Date Modern Atomic Theory continued
GROUND STATE AND EXCITED STATE The lowest state of energy of an electron is called its ground state. When an electron gains energy, it moves to an

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excited state in a higher energy level. Electrons gain energy by absorbing photons. A photon is the smallest unit of light energy.

CHAPTER 4 Atoms SECTION 3 Modern Atomic Theory

Modern Models of the Atom. In the

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modern atomic model, electrons can be found only in certain energy levels, not between levels. Electron location is limited to energy levels. Electrons act like waves. The exact location of an electron cannot be determined.

Orbital: a region in an atom where there is a high probability of finding

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Modern Atomic Theory and Electrons
Chapter 4 Section 3 . Energy Levels !
In Bohr's model, electrons move at a
... Atomic Orbitals ! The electron cloud
represents all the ... 3! 9! 18! 4! 16! 32!
Electron Configurations !

Modern Atomic Theory and Electrons

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Matter is composed of exceedingly small particles called atoms. An atom is the smallest unit of an element that can participate in a chemical change. An element consists of only one type of atom, which has a mass that is characteristic of the element and is the same for all atoms of that element

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*2.3: Modern Atomic Theory and the
Laws That Led to It ...*

3—3 Review and Reinforcement

Modern Atomic Theory Complete the following table. 1. 2. 3. If the statement is true, write “true.” If it is false,

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change the underlined word or words to make the statement true. Write your answer on the line provided. 4.

Moseley discovered that all atoms of an element have the same number of neutrons in ...

Name Date Class 3—3 Review and

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Reinforcement Answers

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and motion. (Section 3.1) Describe the nuclear model of the atom. (Section 3.4) Describe the similarities and differences between solids, liquids, and gases with reference to the particle nature of matter, the degree of motion of the particles, and the degree of attraction between the particles.

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(Section 3.1) Review Skills

Chapter Modern a theory

What did Bohr contribute to modern atomic theory? Bohr contributed a description of the arrangement of electrons in an atom. ... Science 4.3
Chapter 4 section 3. 16 terms.

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gonzalis. 04 - Atomic Structure:

Section 3. 13 terms. PosberghTenafly.

OTHER SETS BY THIS CREATOR.

5.1. 5 terms. jason8876. Chapter V
(first part) 25 terms.

4.3 Flashcards / Quizlet

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Chapter Four [Arrangement of Electrons in Atoms] Chapter 4 :
Arrangement of electrons in atoms.
Taken from the book Modern Chemistry by Holt, Rinehart, and Winston on Chapters 4 and 5, which

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deals with electrons and the periodic table.

*Chapter 4 Arrangement Of Electrons
In Atoms Section 3*

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*Modern Atomic Theory Practice Quiz -
Name_ Class_ Date ...*

Chemistry is based on the modern
atomic theory, which states that all

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Matter is composed of atoms. Atoms themselves are composed of protons, neutrons, and electrons. Each element has its own atomic number, which is equal to the number of protons in its nucleus. Isotopes of an element contain different numbers of neutrons.

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3.1 Atomic Theory - lardbucket

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All matter is composed of atoms •

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Atomic Theory (pages 113-118) . Use this table to answer questions 12 and 13.. Name Chapter 4: Atomic Structure Worksheet .

*Chapter 4 Atomic Structure Worksheet
Answer Key Pdf*

We will investigate that more = in the

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next Section SECTION 3-1 HW: Page 69; 1, 2, & = 3 and Page 87; 1 & 2 .

Page 69. 1. Describe the major contributions of each of the following to the modern theory of the atom:
&nbs= p; a. Democritus = /p> b. = John Dalton. 2= . List the five essential points of Dalton's atomic theory. 3= .

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